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(54) Process for the preparation of a pentenoate ester

(57) A process is described for the preparation of a pentenoate ester by carbonylation of butadiene or butadiene derivate in the presence of carbon monoxide, an alcohol and a catalyst system comprising palladium, a

carboxylic acid and a phosphine ligand in which the phosphine ligand is not a bidentate phosphine ligand in which the bridging group is a bis(η -cyclopentadienyl) coordination group of a transition metal, wherein the carboxylic acid is pentenoic acid.

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Description

The invention relates to a process for the preparation of a pentenoate ester by carbonylation of butadiene or butadiene derivative in the presence of carbon monoxide, an alcohol and a catalyst system comprising palladium, a carboxylic acid and a phosphine ligand in which the phosphine ligand is not a bidentate phosphine ligand in which the bridging group is a bis (η -cyclopentadienyl) coordination group of a transition metal.

A carbonylation reaction according to this invention means every reaction between a nonsaturated substrate, a hydroxy group containing reactant and carbon monoxide.

US-A-5028734 describes a process in which the carbonylation of butadiene is performed with a catalyst system comprising palladium, a substituted or unsubstituted benzoic acid and a multidentate phosphine ligand, such as 1,4-bis (diphenylphosphino)butane.

A disadvantage of the process of US-A-5028734 is that the catalytically active carboxylic acid reacts with the alcohol to form the corresponding, catalytically inactive, ester. The loss of the acid co-catalyst is disadvantageous because the overall catalyst activity will then decrease and a lower selectivity to pentenoate ester will result. In, for example, a continuous process it is important that the catalyst activity remains at a certain high level over a longer period of time. In order to keep the catalyst system of US-A-5028734 active, fresh carboxylic acid must be added to the reaction system which results in a high consumption of carboxylic acid.

Substituted benzoic acid, which acid according to US-A-5028734 is preferably used as the carboxylic acid, is not easily available and therefore expensive. The need to add fresh substituted benzoic acid will make this process economically unattractive for a commercial large scale plant.

An object of the present invention is to provide an economically attractive process for the preparation of pentenoate ester by carbonylation of butadiene.

This object of the invention is achieved in that the carboxylic acid is pentenoic acid.

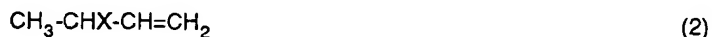
It has been discovered that when pentenoic acid is used as the acid co-catalyst the catalyst activity is comparable to when substituted benzoic acid is used. Advantageously when only pentenoic acid is used as the carboxylic acid, it reacts with the alcohol and the product of the present process (the pentenoate ester) will be formed. Therefore no extra side products, such as, for example a benzoate ester as in the process of US-A-5028734, will be formed in the reaction mixture. This is advantageous because, for example, the separation of the various components in the resulting reaction mixture can be performed more simply.

Another advantage is that fresh pentenoic acid which is needed to compensate the pentenoic acid lost during the carbonylation can be easily obtained, for example, by hydrolysis of the pentenoate ester in a separate step. This manner of obtaining fresh carboxylic acid by hydrolysis of the ester formed is also possible when using the state of the art acids as described in US-A-5028734. The advantage of using pentenoic acid is, however, that the ester which is to be hydrolysed need not be isolated from the pentenoate ester as would be the situation with, for example, a benzoate ester. Therefore a process using pentenoic acid can be performed in an economically attractive manner.

Furthermore, it has been discovered that by adding small amounts of water, fresh pentenoic acid as described above does not have to be supplied. This will be described below in more detail.

The term butadiene derivatives means those compounds which yield pentenoate ester or pentenoic acid as the major product when carbonylated in the process according to the invention. If no statement is made to the contrary, all references to butadiene shall also include butadiene derivatives in this description. It is also possible to carbonylate mixtures of butadiene and butadiene derivatives with the process according to the invention. Although butadiene derivatives can be readily carbonylated in the present process, butadiene is preferred because of its availability. The butadiene can be used in pure form, or in admixture with aliphatic compounds. An exemplary of such admixture is the C₄-cut obtained in a steam cracker process. Such a C₄-cut can comprise butadiene plus 1-butene, 2-butene, and/or isomeric butynes.

Preferred butadiene derivatives are represented by the following general formulae:



wherein X represents an organic group with 1 to 20 carbon atoms or an inorganic group. Examples of suitable organic groups are -OR⁴ or -OC(O)R⁵ groups, in which R⁴ and R⁵ can be, for example, a C₁-C₆ alkyl, C₂-C₆ alkenyl, C₆-C₁₄ aryl, C₇-C₁₄ aralkyl or C₇-C₁₄ alkaryl group. Examples of these groups are methyl, ethyl, n-propyl, isopropyl, n-butyl, tert-butyl, pentyl, cyclopentyl, cyclohexyl, hexyl, propenyl, butenyl, pentenyl, phenyl, naphthyl, benzyl, tosyl. Examples of other suitable organic groups or inorganic groups are -OH, -H₂PO₄, -PR⁶R⁷, -NH-CO-R⁸, -NH₂, and -SR⁹, in which R⁶, R⁷, R⁸ and R⁹ can be the same as defined above for R⁴ and R⁵.

Butadiene derivatives include, for instance, 1-methoxy 2-butene, 3-methoxy 1-butene, 1-ethoxy 2-butene, 3-ethoxy

1-butene, isomeric butenylpentenoate, 1-butene 3-carbonate, 2-butene 1-carbonate, 3-hydroxy 1-butene, and 1-hydroxy 2-butene. Methods of making alkoxy butenes, e.g. methoxy butene, are described in US-A-4343120 and US-A-4590300, the complete disclosures of which are incorporated herein by reference.

The phosphine ligand used in the process according to the invention can be a monodentate or multidentate phosphine ligand, or a mixture thereof.

When a monodentate phosphine ligand is used, the reaction is preferably performed continuously or semi-continuously wherein the amounts of the reactants are substantially maintained during the whole process such that the molar ratio of butadiene to palladium in the reaction mixture is less than 70:1, the molar ratio of carboxylic acid to palladium is greater than 10:1, and the molar ratio of alcohol or water to butadiene is less than 2:1. Preferably the butadiene/palladium ratio is greater than 1:1 and more preferably it is higher than 2:1. It has been found that a high selectivity to pentenoate ester can be obtained when the carbonylation is performed in the above described manner.

To achieve such low butadiene/palladium ratios the butadiene is preferably continuously supplied to the carbonylation reaction at a rate of at most 150 mol butadiene per hour per mol palladium present during the carbonylation. More preferably this rate is less than 80 mol butadiene per hour per mol palladium at a temperature of the carbonylation at or below 160°C.

The amounts of reactants specified herein above should preferably be maintained substantially during the whole process. By preference, the term substantially during the whole process means more than 90% of the process as expressed in residence time.

An additional advantage of the use of monodentate phosphine ligands is that the amount of 2-pentenoate ester in the pentenoate ester mixture formed is lower than when bidentate phosphine ligands, such as described in US-A-5028734, are used in the process according to the invention. A "low amount" means an amount of less than 10% 2-pentenoate ester relative to the total amount of pentenoate esters. This is advantageous when such a mixture is used in the hydroformylation of pentenoate ester to the terminal 5-formylvalerate ester with a rhodium-based catalyst system, such as described, for example, in US-A-5264616, the complete disclosure of which is incorporated herein by reference. The 2-pentenoate ester in the mixture has an adverse effect on the selectivity to 5-formylvalerate ester as is clear from US-A-5264616.

The monodentate phosphine ligand is preferably a compound represented by the following general formula (3):



wherein R¹, R² and R³ each individually represent an optionally substituted organic group. This organic group can be a C₁-C₂₀ alkyl group, a C₂-C₂₀ alkenyl group, a C₆-C₁₈ aryl group or a cyclic group with 4-12 carbon atoms in which the ring of the cyclic group also contains one or more heteroatoms, for example nitrogen. Alkyl groups include, among others, methyl, ethyl, isopropyl, tert-butyl, cyclohexyl or cyclooctyl. An exemplary alkenyl group is butenyl. Exemplary cyclic groups containing heteroatoms include, among others, 6-methyl-2-pyridyl and 4,6-dimethyl-2-pyridyl. By preference, at least one of the organic groups R¹, R² and R³ is a C₆-C₁₆ aryl group and more preferably a C₆-C₁₄ aryl group. Aryl groups include, for instance, naphthyl, phenyl, benzyl, cumenyl, mesityl, tolyl and xylyl. The organic group can be substituted, for example, with halogen atoms, for example Cl, Br or F, or with C₁-C₆ alkyl, C₆-C₁₈ aryl, C₁-C₆ alkoxy, carboxy, carbalkoxy, acyl, trihalogenmethyl, cyano, dialkylamino, sulphonylalkyl or alkanoyloxy groups. Substituents can be groups with electron withdrawing or electron donating properties.

Monodentate phosphine ligands include, for instance, tri-p-tolylphosphine, tri-p-methoxyphenylphosphine, diphenylpentylphosphine or dimethylphenylphosphine. Preferably triphenylphosphine is used because this compound is readily available.

To simplify the process, by preference no other carboxylic acids besides pentenoic acid are present. Other carboxylic acids can react with the alcohol to form their ester products, which would complicate further processing. Fresh pentenoic acid needed to replace the lost acid of the catalyst system can advantageously be prepared in a separate step by hydrolyzing a portion of the pentenoate ester formed in the carbonylation to the pentenoic acid. The pentenoic acid thus obtained can be used in a following carbonylation in which the catalyst system of the previous carbonylation is reused.

Such a separate hydrolysis can, for example, be performed by contacting some of the pentenoate ester with an acid ion exchange resin in the presence of water. Such contacting can be performed in, for example, an on-purpose unit operation (specific designed process step or equipment) or in one of the distillation columns used for separating the pentenoate ester from one of the other components of the reaction mixture leaving the reactor.

It has been found that the addition of water to the carbonylation reaction will result in a stable concentration of pentenoic acid in a continuous process. The amount of water should be sufficient to maintain a catalytically active level of pentenoic acid during the carbonylation. The amount of water needed will depend on the amount of pentenoic acid

being consumed by esterification during the carbonylation. The rate of esterification will depend on the reaction conditions, selected and can be easily determined by analyzing the reaction mixture leaving the reactor.

Preferred multidentate phosphine ligands are represented by the following general formula (4):



wherein n is 2-6, R⁵ is a multivalent (valency equals n) organic bridging group with 2 to 20 carbon atoms not being a divalent bis(η-cyclopentadienyl) coordination group of a transition metal, and R⁴ and R⁶ each individually represent an optionally substituted organic group. By preference, n is 2 in formula (4). Organic groups for R⁴ and R⁶ are the same as described for R¹, R² and R³. The substituents for the organic groups in formula (4) are the same as described for the monodentate phosphine ligands.

R⁵ is preferably an organic group in which the chain of atoms directly connecting the phosphor-atoms, consists of C, S, O, N and/or Si-atoms and preferably only C-atoms.

Divalent organic bridging groups include C₂-C₁₀ alkylidene groups, for example ethylene, trimethylene, tetramethylene, pentamethylene or trans-1,2-cyclobutene; and C₆-C₂₀ divalent arylene groups such as, for example, dinaphthyl or diphenyl.

The bidentate phosphine ligands include, among others, 1,3-bis(diphenylphosphino)propane, 1,4-bis(diphenylphosphino)butane, 2,3-dimethyl-1,4-bis(diphenylphosphino)butane, 1,4-bis(dicyclohexylphosphino)butane, 1,3-bis(dip-tolylphosphino)propane, 1,4-bis(di-p-methoxyphenylphosphino)butane, 2,3-bis(diphenylphosphino)-2-butene, 1,3-bis(diphenylphosphino)-2-oxopropane, 2-methyl-2-(methyl-diphenylphosphino)-1,3-bis(diphenylphosphino)propane, 2,2'-bis(diphenylphosphino)biphenyl, 2,3-bis(diphenylphosphino)naphthalene, 1,2-bis(diphenylphosphino)cyclohexane, 2,2-dimethyl-4,5-bis(diphenylphosphino)dioxolane, 2,3-o-isopropylidene-2,3-dihydroxy-1,4-bis(diphenylphosphino)-butane (DIOP), trans-1,2-bis(di(m-methylphenyl)-phosphinemethyl)cyclobutane, trans-[(bicyclo[2.2.1]heptane-2,3-diyl)bis(methylene)]-bis[diphenylphosphine], trans-[(bicyclo[2.2.2]octane-2,3-diyl)bis(methylene)]-bis[diphenylphosphine], trans-1,2-bis(diphenylphosphino-methyl)cyclobutane (DPMCB), trans-1,2-bis(diphenylphosphinemethyl)trans-3,4-bis(phenyl)cyclobutane and 2,2'-bis(diphenylphosphino)-1,1'-binaphthyl (BINAP).

The molar ratio of phosphine ligand to palladium depends on the specific phosphine ligand used in the process according to the invention. This ratio will preferably be between 1:1 and 20:1. For multidentate phosphine ligands this ratio is preferably between 1:1 and 10:1. When using monodentate phosphine ligands this ratio is preferably greater than 5:1. When this ratio is too high the catalytic effect of the catalyst system is weaker and by-products such as vinyl cyclohexene and high-molecular weight products may form. Both multidentate and monodentate phosphine ligands can be simultaneously present during the carbonylation.

All inert solvents are in principle suitable as an additional solvent, although it is also possible to use an excess of one of the reactants or (by) products in such an amount that a suitable liquid phase is formed. Examples of (by) products are the pentoate esters, C₉-esters and high boiling by-products. The pentoic acid can also be used as solvent. Examples of inert solvents are sulphoxides and sulphones, such as for instance, dimethyl sulphoxide, diisopropyl sulphone; aromatic solvents, such as benzene, toluene, xylene; esters, such as methyl acetate, methyl valerate, pentoate esters and butyrolactone; ketones, such as acetone or methylisobutyl ketone; ethers such as anisole, trioxanone, diphenyl ether and diisopropyl ether; and mixtures of these solvents. Preferably, diphenyl ether is used as additional solvent.

The palladium can be present in the reaction mixture as a heterogeneous palladium compound or as a homogeneous palladium compound. However, homogeneous systems are preferred. Since palladium in situ forms a complex with the bidentate ligand, the choice of the initial Pd compound is in general not critical. Homogeneous palladium compounds include, for instance, palladium salts of, for instance, nitric acid, sulphonic acid, alkane carboxylic acids with not more than 12 carbon atoms or hydrogen halogenides (F, Cl, Br, I). Metallic palladium can also be used. Exemplary homogeneous palladium compounds include PdCl₂, PdBr₂, PdI₂, Na₂PdI₄, K₂PdI₄, PdCl₂(benzonitrile)₂ and bis(allyl)palladium chloride). Another group of suitable halogen-free palladium compounds are palladium complexes such as palladium acetylacetonate (Pd(acac)₂), Pd(II) acetate, palladiumnitrate Pd(NO₃)₂, o-tolyl phosphine palladium, and di-palladium-tris-(dibenzylideneacetone)Pd₂(dba)₃. An exemplary of a heterogeneous palladium compound is a palladium compound on an ion exchanger such as, for example an ion exchanger containing carboxylic acid groups. Ion exchangers containing carboxylic acid groups are commercially available under the brand names Amberlite IRC 50 and Amberlite IRC 84 (Rohm & Haas). Another heterogeneous catalyst is an immobilized phosphine on carrier catalyst, in which the palladium forms a complex with the immobilized phosphine (phosphine being the ligand of the

catalyst system). Carriers include polystyrene, polyacrylamide, or silica.

The palladium concentration in the reaction mixture is preferably as high as possible because the greater will be the rate of reaction per unit of reactor volume. The upper limit for a homogeneous catalyst system will normally be determined by the solubility of palladium in the reaction mixture and will, for example, depend on the specific palladium compound used as discussed above. This upper limit can easily be determined by one skilled in the art. However, the process according to the invention may also be performed with a homogeneous catalyst system in the presence of solid palladium compounds.

The pentenoate ester which is formed during carbonylation of butadiene will actually be a mixture of 2-, 3- and 4-pentenoate esters. The pentenoic acid used in the process according to the invention can also be 2-, 3- or 4-pentenoic acid or mixtures of any two or three of these isomeric pentenoic acids. The 2- and 3-pentenoic acids and esters are present in their cis and trans configuration.

The molar ratio of pentenoic acid to palladium is preferably greater than 10:1 and more preferably greater than 20:1. An upper limit theoretically does not exist. The pentenoic acid can serve as solvent in this reaction. In practice the upper limit will be determined by the choice of a practical palladium concentration. The palladium concentration is preferably as high as possible as explained above. It has been found that virtually none of the pentenoic acid is carbonylated in the process according to the invention.

The alcohol is, for example, an organic compound with 1 to 20 carbon atoms containing one or more OH-groups per molecule. The organic compound can be an aliphatic, cycloaliphatic or aromatic compound. These compounds include, for instance, phenol, cresol, tert-butyl catechol and cyclohexanol. By preference, the alcohol is an aliphatic alcohol in which the aliphatic group R is a straight or branched chain alkyl group. By preference, the alkyl group has 1 to 6 carbon atoms. These aliphatic alcohols can be alkanols represented by the formula ROH of which methanol, ethanol, propanol, isopropanol, butanol, tert-butanol, pentanol and hexanol are exemplary. Most preferably methanol or ethanol is used as the alcohol. Substituted alcohols can also be used such as, for instance, ether-substituted alcohols, of which the methyl ether of ethylene glycol is exemplary.

The molar ratio of alcohol to butadiene can vary between wide limits and generally lies between 10:1 and 1:10. Preferably the alcohol concentration in the reaction mixture is kept as low as possible while the amount of alcohol is at least the stoichiometric amount relative to butadiene. The low concentration of alcohol can be obtained by performing the carbonylation in the presence of large amounts of, for example, an additional solvent or (side-) products formed by the carbonylation reaction. The molar ratio of alcohol to butadiene is preferably at least 1:1, is preferably less than 3:1 and is more preferably less than 1.5:1.

The temperature of the carbonylation is preferably between 25°C and 200°C. More preferably this temperature is between 80°C-160°C. The pressure is not very critical and will generally be between 1 MPa and 20 MPa and preferably be between 2 MPa and 10 MPa.

The carbon monoxide can be used in a pure form or diluted with an inert gas such as, for example nitrogen, rare gases or carbon dioxide. In general more than 5% hydrogen is undesirable, since this can cause hydrogenation of butadiene under the carbonylation conditions. The amount of carbon monoxide is not critical if at least a stoichiometric amount of carbon monoxide relative to butadiene is supplied to the carbonylation reaction.

The carbonylation can be performed batch wise, semi-continuously or continuously. Preferably a continuous manner of operation is used in a commercial large scale process. When a monodentate phosphine ligand is used a continuous process is especially advantageous as explained above. By operating continuously fresh butadiene can be supplied to the catalyst system at a rate comparable to the rate of conversion of butadiene. Preferably alcohol is also supplied continuously to the catalyst system in a preferred ratio to butadiene as described above. In this manner a reasonable conversion of butadiene per mol palladium per hour and a high yield to pentenoate ester can be obtained. An example of such a mode of operation is a series of continuously stirred tank reactors (CSTR) in which the catalyst system, solvent, butadiene and alcohol are fed to the first reactor in the preferred ratios according to the process of the invention. Carbon monoxide is supplied to every reactor. The effluent of the first reactor is fed to the second reactor. Fresh butadiene and alcohol are also fed to the second reactor in the appropriate amounts and ratios. This procedure is repeated for the subsequent reactors. Instead of a series of CSTR's a tube reactor can, for example, be used in which butadiene and alcohol are supplied at intermediate locations along the tube.

The above described continuous process can also be advantageously used when performing the process with a catalyst system comprising multidentate phosphine ligands.

The invention is also directed to a continuous process for the preparation of pentenoate esters by carbonylation of butadiene or a butadiene derivative according to the process of the invention as described above wherein the following steps are performed;

- (a) carbon monoxide, alkanol, the catalyst system and butadiene or a butadiene derivative are continuously fed to a reactor in which the carbonylation takes place,
- (b) separating from the effluent of the reactor unreacted carbon monoxide, unreacted butadiene and unreacted

alcohol in one or more separation steps and returning these reactants to step (a) and isolating the pentenoate ester, (c) returning the remaining mixture of step (b), containing the catalyst system to step (a), (d) hydrolyzing part of the pentenoate ester in a separate step to pentenoic acid and returning the pentenoic acid to step (a).

Preferably a part of the remaining mixture of step (b) is separated from the mixture and led to a drain (purge) in order to prevent a build up of high boiling by-products in the circulating reaction mixture. In general, the purge stream will be reprocessed to retrieve, for example palladium and/or the phosphine ligand.

Step (a) can be performed in several ways such as, for example, in a continuously stirred tank reactor as described above or a bubble column in which the product is simultaneously stripped from the liquid phase.

Separating the carbon monoxide, butadiene, alcohol and the alkyl pentenoate from the reaction mixture in step (c) can be performed in various ways. Generally the carbon monoxide is separated first from the reaction mixture in for example a simple gas - liquid separation unit. The butadiene, alcohol and the pentenoate ester can be separated from the reaction mixture in one step followed by isolating the pentenoate ester from butadiene and alcohol. Preferably the butadiene and alcohol are separated from the reaction mixture in a separate step followed by the isolation of the pentenoate ester from the remaining reaction mixture. Separation of the various compounds can be performed in various ways such as, for example by a simple flash operation or by distillation. The choice as to which unit operation is the most suitable will depend, for example, on the physical properties of the compounds to be separated.

Pentenoic acid or pentenoate ester can be, for example, advantageously used as an intermediate compound in the preparation of ϵ -caprolactam and adipic acid, raw materials for the preparation of nylon-6 and nylon-6,6, respectively.

The invention will be elucidated with the following non-limiting examples. The conversion (conv.), selectivity (sel.) and activity (act.) mentioned in the examples are defined as follows:

$$\text{conv.} = \frac{\text{converted butadiene (mol)}}{\text{initial amount of butadiene (mol)}} * 100\%$$

$$\text{sel. to pentenoate} = \frac{\text{obtained amount of pentenoate (mol)}}{\text{converted amount of butadiene (mol)}} * 100\%$$

$$\text{act.} = \text{converted amount of butadiene (mol) per mol Pd per hour}$$

In the above conversion, selectivity and activity determinations, the term converted butadiene means the amount of butadiene which is reacted to (by)products which cannot react under the carbonylation reaction conditions in any way to the product 2-, 3- and 4-pentenoate. These (by)products are for example, butene, vinylcyclohexene and high boiling products, for example, C₉-heavies (e.g. nonadienoates) and higher boiling products. Excluded from this list are intermediates which can react to pentenoate.

Example I

A 50 ml Parr autoclave, made of Hastelloy C, was filled successively with 0.124 g (0.552 mmol) of Pd(II) acetate, 0.957 g (2.242 mmol) of 1,4-bis(diphenylphosphino)butane, 0.784 g (7.84 mmol) of 3-pentenoic acid and 10.13 g of diphenyl ether as a solvent. The autoclave was closed and purged three times with 4.0 MPa carbon monoxide. Subsequently, under a pressure of 1.0 MPa of CO while stirring at a speed of 1250 rpm, a mixture of 3.65 g (114 mmol) of methanol, 5.805 g (108 mmol) of butadiene and 0.480 g of nonane (internal standard for GC product analysis) was injected under pressure from an injection vessel into the autoclave. The temperature of the reaction mixture was raised to 140°C at a CO pressure of 6.0 MPa. After 2.0 hours the reaction was stopped and the butadiene and the reaction products were analyzed by gas chromatographic methods.

The conversion was 75%. Selectivity to methyl pentenoates was 79%, and the activity was 58 hr⁻¹. 46% of the 3-pentenoic acid was converted to its methyl ester.

Comparative Experiment A

Example I was repeated using 2,4,6-trimethyl benzoic acid instead of 3-pentenoic acid, in 7 molar equivalents per Pd. After 3.8 hours the reaction was stopped.

The conversion was 86%. Selectivity to methyl pentenoates was 87%, and the activity was 39 hr⁻¹. 23% of the trimethyl benzoic acid was converted into its methyl ester.

Example II

A 150 ml Parr autoclave, made of Hastelloy C, was filled successively with 0.387g (1.73 mmol) of Pd(II) acetate, 2.86 g (6.70 mmol) of 1,4-bis(diphenylphosphino)butane, 1.20 g (12 mmol) of 3-pentenoic acid, and 0.56 g of nonane (internal standard for GC product analysis) and 32.40 g of diphenyl ether as a solvent. The autoclave was closed and purged three times with 4.0 MPa carbon monoxide. Subsequently the temperature of the reaction mixture was raised to 140°C at a CO pressure of 5.0 MPa, after which the butadiene, methanol and 3-pentenoic acid supplies were started at constant rates of 130, 171 and 7.8 mmol per hour respectively, using Gilson model 302 pumps. After 2.8 hours the reaction was stopped and the butadiene and the reaction products analyzed by gas chromatographic methods.

The conversion was 79%, and the selectivity to methyl pentenoates was 82%. The activity was 50 hr⁻¹, and 70% of the 3-pentenoic acid was converted to its methyl ester.

Example III

Example II was repeated, using triphenylphosphine as the ligand at a concentration of 10 equivalents per Pd, and adding 10 equivalents of 3-pentenoic acid prior to butadiene and methanol supply at rates of 40 and 48 mole/mole Pd/hr respectively. Together with methanol 4 equivalents of 3-pentenoic acid per palladium per hour were fed to the autoclave. After 5 hours the reaction was stopped.

The conversion was 81%, and the selectivity to methyl pentenoates was 88%. The activity was 29 hr⁻¹. 44% of the total amount of 3-pentenoic acid was converted to methyl-3-pentenoate.

Comparative Experiment B

A 50 ml Parr autoclave, made of Hastelloy C, was filled successively with 0.05 g (0.22 mmol) of Pd(II) acetate, 0.61 g (2.32 mmol) of triphenyl phosphine, 0.26 g (1.6 mmol) of 2,4,6-trimethyl benzoic acid and 27.0 g of diphenyl ether as a solvent. The autoclave was closed and purged three times with 4.0 MPa carbon monoxide. Subsequently, under a pressure of 1.0 MPa of CO, while stirring at a speed of 1250 rpm, a mixture of 3.76 g (118 mmol) of methanol, 5.14 g (95.3 mmol) of butadiene and 0.45 g of nonane (internal standard for GC product analysis) was injected under pressure from an injection vessel into the autoclave. The temperature of the reaction mixture was raised to 150°C at a CO pressure of 6.0 MPa. After 5.0 hours the reaction was stopped and the butadiene and the reaction products were analyzed by gas chromatographic methods.

The conversion was 85%. Selectivity to methyl pentenoates was 21%, and the activity was 16 (hr⁻¹). 9% of the trimethyl benzoic acid was converted to its methyl ester.

Example IV

A 150 ml Parr autoclave, made of Hastelloy C, was filled successively with 0.43 g (1.94 mmol) palladium acetate, 5.0 g (19 mmol) triphenyl phosphine, 8.5 g (85 mmol) 3-pentenoic acid and 0.80 g nonane (internal standard for GC analysis). The autoclave was closed and purged three times with 4.0 MPa carbon monoxide. Subsequently the temperature of the reaction mixture was raised to 140°C at a CO pressure of 4.0 MPa, and thereafter butadiene was supplied at 83 mmol per hour, and methanol at 83 mmol per hour. After one hour, 2.3 g (126 mmol) of water was injected into the autoclave under pressure. The pressure was raised to 6.0 MPa, and the butadiene and methanol supply continued at the same rates as during the first hour. After an additional two hours the reaction was stopped, and the contents of the autoclave were analyzed by gas-chromatographic methods.

The conversion was 80%. Selectivity to methyl pentenoates was 82%. Only 5% of 3-pentenoic acid was converted to methyl-3-pentenoate.

Example V

Example IV was repeated, except that water was continuously supplied together with methanol, at rates of 54 and 78 mmol per hour respectively. The butadiene supply was 78 mmol per hour. After four hours the reaction was stopped.

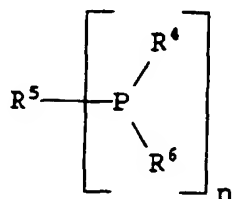
The conversion was 75%, and the selectivity to methyl pentenoates 79%. No (net) esterification of 3-pentenoic acid was observed.

Claims

1. A process for the preparation of a pentenoate ester by carbonylation of butadiene or a butadiene derivative in the

presence of carbon monoxide, an alcohol and a catalyst system comprising palladium, a carboxylic acid and a phosphine ligand in which the phosphine ligand is not a bidentate phosphine ligand in which the bridging group is a bis(η -cyclopentadienyl) coordination group of a transition metal, characterized in that the carboxylic acid is pentenoic acid.

2. A process according to claim 1, characterized in that the molar ratio of pentenoic acid to palladium is greater than 10:1.
3. A process according to claims 1-2, characterized in that the molar ratio of alcohol to butadiene or butadiene derivative is between 1:1 and 3:1.
4. A process according to claims 1-3, characterized in that the alcohol is methanol or ethanol.
5. A process according to claims 1-4, characterized in that a part of the pentenoic ester is hydrolyzed to pentenoic acid, and the thus obtained pentenoic acid is recycled to the carbonylation reaction.
6. A process according to claims 1-5, characterized in that an amount of water is present during the carbonylation in an amount sufficient to maintain a catalytically active level of pentenoic acid.
7. A process according to claims 1-6, characterized in that the phosphine ligand is a monodentate phosphine ligand, the carbonylation is performed continuously or semi-continuously, and the amounts of the reactants are substantially maintained during the carbonylation such that the molar ratio of butadiene or butadiene derivative to palladium is less than 70:1, and the molar ratio of alcohol to butadiene or butadiene derivative is less than 2:1.
8. A process according to claims 1-7, characterized in that the phosphine ligand is a multidentate phosphine ligand with the formula



(4)

wherein n is 2-6, R^5 is a multivalent organic bridging group with 2 to 20 carbon atoms in which the chain of atoms directly connecting the phosphor atoms consists of C, S, O, N and/or Si-atoms, and R^4 and R^6 each individually represent an optionally substituted organic group.

9. A process for preparing a pentenoate ester comprising conducting the carbonylation of at least one member selected from the group consisting of butadiene and a butadiene derivative in a reaction mixture in the presence of carbon monoxide, an alcohol selected from the group consisting of methanol and ethanol, and a catalyst system, wherein said catalyst system is the combination of palladium, a phosphine ligand, and pentenoic acid, characterized in that in said carbonylation the molar ratio of pentenoic acid to palladium is greater than 10:1, the molar ratio of said alcohol to butadiene or butadiene derivative is between 1:1 and 3:1.
10. A process according to claim 9, characterized in that
 - (a) carbon monoxide, the alcohol, the catalyst system and butadiene or a butadiene derivative are continuously fed to a reactor in which the carbonylation takes place,
 - (b) unreacted carbon monoxide, unreacted butadiene and unreacted alcohol are separated from the effluent of the reactor in one or more separation steps and these reactants are returned to step (a) and the pentenoate ester is isolated,
 - (c) the remaining mixture of step (b), containing the catalyst system is returned, to step (a)
 - (d) a part of the pentenoate ester is in a separate step hydrolyzed to pentenoic acid and returned to step (a).



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EUROPEAN SEARCH REPORT

Application Number
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DOCUMENTS CONSIDERED TO BE RELEVANT			
Category	Citation of document with indication, where appropriate, of relevant passages	Relevant to claim	CLASSIFICATION OF THE APPLICATION (Int.Cl.6)
A	EP-A-0 271 145 (SHELL INTERNATIONALE RESEARCH MAATSCHAPPIJ B.V.) * page 3, line 32 - page 4, line 18 * * page 6, line 6 - line 16 * * page 7, line 29 - line 33 * * page 8, line 34 - page 9, line 16 * * page 11 - page 14; examples 1-3,5 * * page 15 - page 18; claims *	1	C07C67/38 C07C69/533
A	EP-A-0 577 205 (SHELL INTERNATIONALE RESEARCH MAATSCHAPPIJ B.V.) * page 3, line 13 - line 36 * * page 4, line 51 - page 5, line 15 * * page 4 - page 5; examples 1-10 * * page 6 - page 7; claims *	1	
A	EP-A-0 273 489 (SHELL INTERNATIONALE RESEARCH MAATSCHAPPIJ B.V.) * page 3, line 17 - line 40 * * page 4, line 32 - line 50 * * page 5, line 11 - line 22 * * page 7; examples 5,7 * * page 8 - page 9; claims *	1	
D	& US-A-5 028 734		
			TECHNICAL FIELDS SEARCHED (Int.Cl.6)
			C07C
The present search report has been drawn up for all claims			
Place of search THE HAGUE		Date of completion of the search 9 May 1996	Examiner Kinzing, J
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